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ChemActivity 50 The Electrochemical Cell $293 \text{ Cu}^{2+}(1 \text{ M}) + 2\text{e}^- \text{ Cu}(s)$ Simultaneously, at the Pt electrode (anode), the following reaction takes place: $\text{H}_2(\text{g}; 1 \text{ atm}) \rightarrow 2\text{H}^+(1 \text{ M}) + 2\text{e}^-$ The experimental voltage, E° , is 0.34 V. Critical Thinking Questions 11.

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9. Which is more concentrated with respect to sodium ions, 0.50 M NaCl or 0.30 M Na₂SO₄? 10. Which CTQ was easier to answer, 8 or 9? Why? 182 ChemActivity 32 Molarity Information Some molecules do not dissociate into ions when dissolved in water. Sugars (glucose, sucrose, dextrose, etc.) and alcohols are examples.

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Chem 124 Homework

ChemActivity Colligative Properties 3. Calculate the mass of the solvent added to the solution in model 1 (in kg). 4. What is the molality of the KCl solution in model 1? 5. Why are your answers to CTQ 3 and 4 different? 6. Which has the greater molality, 50.0 g of NaCl dissolved in 100.0 ml of water or 50.0 g of NaCl dissolved in 1000 ml of ...

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