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Note as well that since the system is isolated, no heat has been exchanged with the surroundings ($q = 0$). The first law of thermodynamics confirms that there has been no change in the system's internal energy as a result of this process.

$$\Delta U = q + w$$

\tag{First Law of Thermodynamics}

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$$\Delta G = 0 + 0 = 0$$

$$\text{\label{Eq3}\end{align}\}$$

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