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Modern Chemistry Chapter 14 Acids & Bases Section 1 Properties of Acids & Bases Aqueous solutions of acids taste sour. Acids change the color of acid-base indicators. Some acids react with active metals and release hydrogen gas. Acids react with bases to produce salts and water. Acids conduct electricity (electrolytes).

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CHAPTER 14 REVIEW Acids and Bases
SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1: $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{SO}_3^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$
stage 2: $\text{HSO}_3^-(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{SO}_3^+(\text{aq}) + \text{SO}_3^{2-}(\text{aq})$

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(aq) SO₃ (aq) b.

14 Acids and Bases

Modern Chemistry 121 Acids and Bases
CHAPTER 14 REVIEW Acids and Bases
SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Answer the following questions according to the Brønsted-Lowry definitions of acids and bases:
_____ a.

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Chapter 14 Review, Section 2: Acids and Bases Lewis acids and bases, Brønsted-Lowry acids and bases, ionization, and more are covered by this chemistry handout. It serves as a review of a specific textbook section, but will also serve as a nice review for any general chemistry

Chapter 14 Review Acids Bases Section 1 Modern Chemistry

Chemistry: Chapter 14 (Section 14.1)

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CHAPTER 14 REVIEW Ions in Aqueous Solutions and Colligative Properties
SECTION 14-1 SHORT ANSWER Answer the following questions in the space provided. 1. Use the guidelines in Table 14-1 on page 427 of the text to predict the solubility of the following compounds in water: a. magnesium nitrate b. barium sulfate c. calcium carbonate d. ammonium phosphate

14 Ions in Aqueous Solutions and Colligative Properties

Read Book Modern Chemistry Worksheet Answers Chapter 14 Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation?

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compounds, or mixtures based on

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questions in the space provided.

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of the textbook. Pressure Laws. standard
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Modern Chemistry 6 The Periodic Law
CHAPTER 5 REVIEW The Periodic Law

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SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. ____ When an electron is added to a neutral atom, energy is (a) always absorbed.

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