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1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient x in this equation? 4

b. What is the molar mass of C_3H_4 ? 40.07 g/mol

c. What is the mole ratio ...

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3 PROBLEMS Write the

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provided. 1. _____ The

actual yield of a

reaction is 22 g and

the theoretical yield is

25 g. Calculate the

percentage yield. 2.

6.0 mol of N_2 are

mixed with 12.0 mol of

H

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Much of our knowledge of chemistry is based on the careful quantitative analysis of substances involved in chemical reactions.

Composition stoichiometry deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical

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reaction.

Section 1 Answers

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following questions in
the space provided. 1.

_____ When an electron
is added to a neutral
atom, energy is (a)
always absorbed.

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Reactions SECTION 82
SHORT ANSWER

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